

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property Organization
International Bureau



(43) International Publication Date
1 August 2002 (01.08.2002)

PCT

(10) International Publication Number
WO 02/059066 A1

(51) International Patent Classification⁷: C07C 11/04,
11/107, 6/04, 5/25

(21) International Application Number: PCT/US02/02352

(22) International Filing Date: 24 January 2002 (24.01.2002)

(25) Filing Language: English

(26) Publication Language: English

(30) Priority Data:
60/263,924 25 January 2001 (25.01.2001) US

(71) Applicant: ABB LUMMUS GLOBAL INC. [US/US];
1515 Broad Street, Bloomfield, NJ 07003 (US).

(72) Inventor: GARTSIDE, Robert, J.; 167 Colonial Road,
Summit, NJ 07901 (US).

(74) Agents: LILLIE, Raymond et al.; Carella, Byrne, Bain,
Gilfillan, Cecchi, Stewart & Olstein, 6 Becker Farm Road,
Roseland, NJ 07068 (US).

(81) Designated States (*national*): AE, AG, AL, AM, AT, AU,
AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CU,
CZ, DE, DK, DM, DZ, EE, ES, FI, GB, GD, GE, GH, GM,
HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK,
LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX,
MZ, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL,
TJ, TM, TR, TT, TZ, UA, UG, UZ, VN, YU, ZA, ZW.

(84) Designated States (*regional*): ARIPO patent (GH, GM,
KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW),
Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM),
European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR,
GB, GR, IE, IT, LU, MC, NL, PT, SE, TR), OAPI patent
(BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR,
NE, SN, TD, TG).

Published:

- with international search report
- before the expiration of the time limit for amending the
claims and to be republished in the event of receipt of
amendments

*For two-letter codes and other abbreviations, refer to the "Guid-
ance Notes on Codes and Abbreviations" appearing at the begin-
ning of each regular issue of the PCT Gazette.*



WO 02/059066 A1

(54) Title: PROCESS FOR THE PRODUCTION OF LINEAR ALPHA OLEFINS AND ETHYLENE

(57) Abstract: Ethylene and hexene-1 are produced from butene-1 by metathesis of butene-1 and isomerization of the hexene-3 produced therein to hexene-1. The initial starting material is a mixed butene stream wherein butene-1 is isomerized to butene-2 with isobutylene being separated therefrom, followed by isomerization of butene-2 to butene-1, with the butene-1 being the feed to the metathesis.

Process for the Production of Linear Alpha Olefins and Ethylene

This application claims priority based on, and is a continuation – in-part of, application Serial No. 60/263,924, filed January 25, 2001, the contents of which are incorporated herein by reference in their entirety.

This invention relates to the production of linear alpha olefins. More particularly, this invention relates to the production of linear alpha olefins having a higher carbon number from alpha olefins having a lower carbon number. More particularly, this invention relates to the production of ethylene and hexene-1.

Linear alpha olefins currently are produced by the selective oligomerization of ethylene. In general, the oligomerization is conducted in the presence of a catalyst, such as an alkylated metal catalyst. Long residence times are used to produce hydrocarbon chains of varying lengths. The olefin double bond remains at the alpha position as each ethylene molecule is added. Ethylene oligomerization produces a wide spectrum of linear alpha olefin products. Extensive fractionation is required to separate the alpha olefins having different carbon numbers.

In accordance with an aspect of the present invention, a linear alpha-olefin having a first number of carbon atoms is subjected to a metathesis reaction under conditions and in the

presence of a catalyst that minimizes or eliminates skeletal and/or double bond isomerization to produce a reaction product that includes ethylene and a linear internal olefin (LIO) that has a number of carbon atoms greater than the first number of carbon atoms. Specifically, the feed linear alpha olefin with carbon number n is metathesized to ethylene and linear internal olefin with carbon number $2n-2$. The resulting linear internal olefin may then be isomerized to produce a linear alpha olefin.

In accordance with an aspect of the present invention, butene-1 is subjected to a metathesis reaction under conditions and in the presence of a metathesis catalyst that minimizes or eliminates skeletal and double bond isomerization to produce a reaction product that includes ethylene and hexene-3. The hexene-3 is then isomerized to hexene-1.

In the metathesis reaction, the catalyst and reaction conditions are such as to minimize isomerization of the butene-1 starting material.

The catalyst used in this aspect for the metathesis reaction may be a supported or unsupported catalyst and the overall catalyst is one that has a minimized amount of both acidic and basic sites.

A catalyst that has a reduced amount of both acidic and basic sites (preferably essentially no acidic and basic sites) improves the selectivity of the metathesis of the lower carbon number alpha-olefin (1-butene) with itself to form ethylene and a linear internal olefin of higher carbon number (3-hexene) by minimizing isomerization.

Metathesis reactions can be characterized as either fully productive, half productive, or non-productive depending upon the symmetry of the molecule with respect to the double bond. Two dissimilar olefins will react in a "fully productive" manner. An example is the reaction of 1-butene with 2-butene. The double bond is in a different position in the respective molecule and these will react rapidly with one another. Two identical molecules will react in either a half productive or non-productive manner. If for example 1-butene reacts with itself, because the double bond is in the same position within each molecule, it will react at a rate considerably

slower than the rate of fully productive reactions. If the feed molecules are identical and symmetric around the double bond (for example 2-butene reacting with itself), then no reaction will occur and the system is defined as non-productive.

In many metathesis reaction systems, isomerization activity is incorporated to increase the rates of reaction. By allowing for a shifting double bond, half or non-productive reactions can be reduced. Typical isomerization catalysts include basic metal oxides or promoted zeolites.

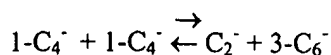
In one preferred embodiment, the catalyst or catalyst mixture contains essentially no magnesium oxide in that magnesium oxide catalyst promotes isomerization. Thus, for example, a preferred catalyst (supported or not supported) is a Group VI B or Group VII B metal oxide such as tungsten oxide, molybdenum oxide, or rhenium oxide, with tungsten oxide being particularly preferred.

If a support is used, such support has a minimized amount of both acidic and basic sites and preferably essentially no acid and essentially no basic sites. Preferred supports are silica or zirconia oxide.

Although the catalyst may include a promoter to reduce acidity; for example, an alkali metal (sodium, potassium or lithium), cesium, a rare earth, etc., in a preferred embodiment, the catalyst does not include a promoter.

Generally reaction conditions that tend to favor the primary reaction and discourage subsequent reactions are preferred. Thus a lower pressure and shorter residence times tend to minimize the isomerization reaction.

In accordance with the present invention, a linear alpha olefin with carbon number n , such as butene-1, is subjected to auto-metathesis with itself (half productive reaction) in the presence of a catalyst and under conditions that minimize isomerization of the linear olefin to produce a reaction product that includes ethylene and a linear alpha olefin of carbon number $2n-2$, such as hexene-3. The reaction is:



The metathesis reaction, in accordance with the invention may be effected at a temperature from about 50°C to 600°C, preferably from about 200°C to 350°C, at a weight hourly space velocity (WHSV) of from about 3 to about 200, preferably from about 6 to about 40, and at a pressure of from about 10 psig to about 600 psig, preferably from about 30 psig to about 100 psig.

The butene-1 starting material may be a pure or impure feed,. The portion of the feed that contains reactive C₄ olefins is preferably at least 90% and more preferably at least 95% butene-1. Non reactive components , for example normal butane, may dilute the olefins in the feedstock. Most preferably, the butene-1 starting material is a pure feed i.e. the butene-1 is present in an amount of at least 99%.

In one embodiment, a mixed butene feed is pretreated to provide a butene-1 feed for the metathesis reaction. For example, the mixed butene feed may contain butene-1, butene-2 and isobutylene. Such a feed may be obtained from a steam cracker. In the event that the mixed butene stream includes butadiene such butadiene is removed by hydrogenation or extraction.

In one embodiment, the mixed butenes feed following or in conjunction with butadiene hydrogenation is subjected to hydroisomerization conditions to convert butene-1 to butene-2, with isobutylene being separated from a butene-2 stream by fractionation. The butene-2 stream is then isomerized back to butene-1 in a subsequent step for use as feed to the metathesis portion of the process

The hydro isomerization of butene-1 to butene-2 and separation may be accomplished in separate steps or may be combined into a single step. In one embodiment butene-1 is isomerized to butene-2 and isobutylene is separated from butene-2 , in a catalytic distillation tower. In another embodiment, the isomerization of the butene-1 to butene-2 can also be accomplished by using a fixed bed isomerization reactor or by adjusting the conditions of the butadiene

hydrogenation reactor followed by a fractionation tower to produce a butene-2 stream and an isobutylene product.

The isomerization of 1-butene to 2-butene is favored by low temperatures. This reaction is carried out typically in liquid phase either in a catalytic distillation tower or fixed bed reactor as described above.

The subsequent isomerization of 2-butene to 1-butene is favored by higher temperature and preferably is carried out in a vapor fixed bed phase reactor. Alternately, the conversion of butene-2 to butene-1 and the purification of the butene-1 may be accomplished by catalytic distillation. The production of 1-butene from hydrocarbons fed to a steam cracker is described further in U.S. Patent No. 5,087,780, the contents of which are incorporated herein by reference.

As a further alternative, a mixed butene stream is reacted with for example methanol to convert isobutylene to methyl tertiary butyl ether (MTBE) as known in the art. The MTBE product is removed from the butenes stream. The resultant mixed normal butenes stream (butene-1 and butene-2) is then processed in a similar manner to the butene-2 stream above to produce an essentially pure butene-1 stream for feed to the metathesis reactor.

Thus, in accordance with an aspect of the present invention a mixed C4 stream is converted to ethylene and hexene-1 by the steps of a. hydrogenating the butadiene to 1 and 2 butenes, b. converting butene-1 to butene-2 and separating isobutylene therefrom, c. isomerization of butene-2 to produce butene-1, d. separation of butene 1 from butene 2 to produce an essentially pure butene 1 stream, e. autometathesis of butene-1 under non-isomerizing conditions to produce ethylene and hexene-3, and f. isomerizing hexene-3 to hexene-1.

The effluent from the metathesis reactor is passed to a series of fractionation towers, where ethylene is removed, unconverted butenes are separated for recycle, and the hexene fraction is removed from the bottoms. The bottoms stream, which is primarily 3-hexene, is isomerized to 1-hexene and the 1-hexene purified by fractionation.. In a preferred embodiment,

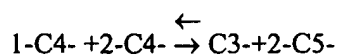
isomers of 2-hexene and 3 hexene are reacted essentially to extinction by recycle. The isomerization reaction schemes for hexene are essentially the same as described above for conversion of 2-butene to 1-butene. The overhead product is 1-hexene.

The hexene-3 produced in the autometathesis is isomerized in the presence of a suitable catalyst. The reaction can take place either in the liquid phase or the vapor phase. For liquid phase reaction, such catalysts include, but are not limited to, palladium and platinum catalysts. The catalyst may be supported on a suitable support material, such as alumina, for example. The reaction occurs in the presence of small amounts of hydrogen (hydroisomerization).

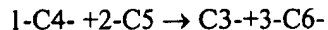
For vapor phase reaction, such catalysts include, but are not limited to, basic metal oxides, including magnesium oxide. In accordance with the invention, the metathesis reaction of the butene-1 occurs first, followed by the isomerization of the hexene-3 product from that reaction. In this case, no hydrogen is required. The equilibrium of mixed hexenes to 1-hexene is favored by higher temperatures hence the vapor phase reaction occurring at higher temperatures is preferred.

The hexene-3 may be isomerized to hexene-1 at a WHSV of from about 3 to about 200, preferably from about 10 to about 60, and at a pressure of from about 2 bar to about 40 bar, preferably from about 3 bar to about 10 bar, and a temperature from about 40°C to 300°C, preferably from about 60°C to 150°C. for liquid phase and 300 to 400 C for vapor phase reaction.

Although in accordance with the present invention, the metathesis reaction is effected with a catalyst and under conditions that minimize isomerization, some isomerization occurs. As a result of the isomerization, the internal olefin of carbon number n, such as butene-2, is produced which reacts with the feed linear alpha olefin of carbon number n, such as butene-1, to produce propylene and a non-selective linear internal olefin of carbon number 2n-3, such as pentene-2, via the reaction:



The propylene may be recovered as reaction product; however, in many cases, the internal olefin of carbon number $2n-3$ is less valuable. In accordance with an aspect of the invention, pentene-2 is recycled to the metathesis reaction. The pentene-2 reacts with butene-1 to produce propylene and hexene-3 via the reaction:



For example, in the metathesis of butene-1 under minimum isomerization conditions, there is about a 35 mol % selectivity of each of ethylene and hexene and 15 mol % of each of propylene and 2-pentene. This gives a 20% wt. selectivity to $2-C_5H_{10}$. If the metathesis feed in addition to the 1-butene includes a stream of 2-pentene, such that the 2-pentene concentration in the feed is 10%, the net pentene selectivity goes to essentially 0 and the hexene selectivity increases from 35 to 46% molar or over 60% by weight. In this manner, hexene selectivity is increased.

In addition to participating in the above reaction, the presence of 2-pentene suppresses the formation of additional $2C_5$ by limiting the non-selective reaction of 1-butene with 2-butene due to equilibrium and ultimately limiting the isomerization of 1-butene to 2-butene because 2-butene formed is not reacting away, thus creating an additional equilibrium limitation.

In accordance with the present invention, the metathesis reaction of 1-butene is effected at conditions and with a catalyst that minimizes isomerization in order to increase selectivity to hexene-3 and ethylene. In particular, such catalyst and conditions are selected in order to achieve a weight selectivity to hexene-3 of at least 40% and preferably at least 50% or greater from 1-butene (without recycle) and still higher values when incorporating pentene recycle.

The invention now will be described with respect to the following example; however, the scope of the present invention is not intended to be limited thereby.

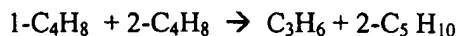
EXAMPLE 1

In this example, a catalyst consisting of WO₃ on a silica support is loaded into a reactor either alone, or admixed with MgO prior to loading. Pure 1-butene is passed over the catalyst at a WHSV of 13 and at various pressures as shown in Table 1 below. The reaction temperature is 650°F (343 C) over the catalyst. Selectivity is calculated by dividing the weight of each product by the weight loss of 1-butene converted to products. The selectivities to various components are given in Table 1 below.

Table 1**Weight percent Selectivity of 1-butene conversion to products**

Residence Time	3/1 MgO/WO3		1/1 MgO/WO3		Pure WO3 12 sec	Pure WO3 4 sec	Pure WO3 2 sec
	8 sec		12 sec				
Pressure	275 psi	275 psi	450 psi	450 psi	450 psi	150 psi	75 psi
C2-	2.5	2.5	2.1	2.4	13.9	17.8	17.4
C3-	26.3	22.9	23.3	23.9	14.9	9.2	9.1
2-C4-	31.2	31.6	32.8	30.4	2.8	1.0	0.6
2-C5-	19.2	18.1	24.8	25.0	16.9	11.0	11.2
1-C5-	2.5	2.5	1.8	2.0	5.8	3.6	1.8
C6	8.4	12.1	10.1	10.8	36.2	50.5	54.1
C7	2.3	5.5	2.1	2.3	4.8	3.1	1.9
C8+	3.7	5.5	1.0	1.0	2.0	1.3	3.9

As shown in Table 1 above, both the 3/1 and the 1/1 ratio mixtures of MgO and WO₃-SiO₂ catalyst show low weight selectivity to ethylene and C₆ alkenes. The isomerization activity of the MgO effectively converts some portion of the 1 butene to 2-butene and the metathesis reaction of:



occurs limiting the auto-metathesis reaction of 1-C₄H₈ to form C₂H₄ + 3-C₆H₁₂.

When a pure WO₃-SiO₂ catalyst is employed, the selectivity for hexene increased from 8.4% - 12.1% to as high as 54.1% hexene. The theoretical weight selectivities for the pure

autometathesis reaction itself (no isomerization) are 25 wt. % ethylene and 75 wt. % hexene. Thus, reducing the isomerization activity of the catalyst system improved selectivity to hexene.

EXAMPLE 2

In this example, a feed containing pure 1-butene, and feeds containing 1-butene mixed with 2-butene or isobutylene are reacted in a reactor in the presence of a catalyst consisting of WO_3 and a silica support at a temperature of 600°F or 650°F, and a pressure of 75 psig. These conditions and catalyst are preferred to minimize isomerization reactions. Selectivity is calculated in wt. % or mol. % by dividing the weight and number of moles of each product by the weight loss and molar loss of 1-butene converted to products. The selectivities to various components are given in Table 2 below.

Table 2
Feed (wt. %)

	<u>A</u>	<u>B</u>	<u>C</u>
1-C ₄	100	75	95
2-C ₄	0	25	0
i-C ₄	0	0	5
Temp. (°F)	650	600	600
Pressure (psig)	75	75	75
<u>Conversion (%)</u>			
1-C ₄	47	52	56
2-C ₄	N/A	66	N/A
i-C ₄	N/A	N/A	31
<u>Selectivity (mol %)</u>			
C ₂ H ₄	37	15	31
3-C ₆ H ₁₂	38	19	28
C ₃ H ₆	12	34	19
2-C ₅ H ₁₀	11	29	18
<u>Others</u>	2	3	4

Selectivity (wt.%)

C ₂ H ₄	18.5	7.5	15.5
3-C ₆ H ₁₂	57	28.5	42
C ₃ H ₆	9	25.5	14
2-C ₅ H ₁₀	14	36.0	22.5
Others	1.5	2.5	6

The above results show that when the feed containing 1-butene is a feed of pure 1-butene with no other components, there is increased selectivity to 3-hexene, as compared with feeds containing 2-butene or isobutylene along with 1-butene.

EXAMPLE 3

In this example, a feed containing essentially pure 1-butene (95 % 1-butene and 5 % iso-butene) and a feed containing 1-butene, iso-butene and 10 % 2-pentene are reacted in a reactor in the presence of a catalyst consisting of WO₃ and a silica support at a temperature of 650°F, and a pressure of 75 psig. These conditions and catalyst are preferred to minimize isomerization reactions. Selectivity is calculated in wt. % and mol. % by dividing the weight and number of moles of each product by the weight loss and molar loss of 1-butene converted to products. The selectivities to various components are given in Table 3 below.

Table 3
Feed (wt. %)

	<u>A</u>	<u>B</u>
1-C ₄	95	87
I-C ₄	5	3
2-C ₅	0	10
Temp. (°F)	650	650
Pressure (psig)	75	75

<u>Conversion (%)</u>		
1-C ₄	56	47
2-C ₅	N/A	26
i-C ₄	31	29
<u>Selectivity (mol %)</u>		
C ₂ H ₄	31	34
3-C ₆ H ₁₂	28	43
C ₃ H ₆	19	17
2-C ₅ H ₁₀	18	1.6
<u>Others</u>	4	4.4
<u>Selectivity (wt.%)</u>		
C ₂ H ₄	15.5	16.9
3-C ₆ H ₁₂	42	64.1
C ₃ H ₆	14	12.7
2-C ₅ H ₁₀	22.5	2.0
Others	6	4.3

The above results show that when the 2-pentene produced by the non-selective isomerization of the low isomerization catalyst is recycled, there is increased selectivity to 3-hexene, as compared with feed where the 2-pentene is not recycled to the reactor. If the base feedstock was a pure 1-butene stream, the increase in hexene selectivity would be even greater.

EXAMPLE 4

In this example, feeds containing essentially pure 1-butene (99.9% 1-butene), and feeds containing 1-butene and 2-pentene are reacted in a reactor in the presence of Catalyst A or Catalyst B at a temperature of 650°F and a pressure of 75 psig. These conditions are preferred to minimize isomerization reactions. Catalyst A consists of WO₃ and a silica support, and Catalyst B consists of WO₃ and a special chromatographic grade silica support. Catalyst A uses a silica support that contains 2,000 ppm sulfur, thus creating acidic reaction sites. Catalyst B uses a silica support where the sulfur has been reduced to less than 100 ppm. Both catalysts have a low

isomerization activity, and the isomerization activity of Catalyst B is lower than that of Catalyst A. Selectivity is calculated in wt. % and mol% by dividing the weight and number of moles of each product by the weight loss and molar loss of 1-butene converted to products. The selectivities to various components are given in Table 4 below.

Table 4
Feed (wt. %)

Catalyst	<u>A</u>	<u>A</u>	<u>B</u>	<u>B</u>
1-C ₄	99.9	87.6	99.9	91.3
2-C ₅	0	10.2	0	8.7
Temp. (°F)	650	650	650	650
Pressure (psig)	75	75	75	75
<u>Conversion (%)</u>				
1-C ₄	56	45.4	41	38
2-C ₅	<u>Net Production</u>	26.6	<u>Net Production</u>	32.5
i-C ₄	N/A	24.4	N/A	N/A
<u>Selectivity (mol%)</u>				
C ₂ H ₄	31	35.5	44.2	38.4
3-C ₆ H ₁₂	28	46.2	45.8	47.8
C ₃ H ₆	19	15.5	5.2	11.6
2-C ₅ H ₁₀	18	<u>Net Conversion</u>	4.4	<u>Net Conversion</u>
<u>Others</u>	4	2.8	0.4	2.2
<u>Selectivity (wt. %)</u>				
C ₂ H ₄	15.5	17.8	22.1	19.2
3-C ₆ H ₁₂	42	69.3	68.6	71.7
C ₃ H ₆	14	11.6	3.8	8.7
2-C ₅ H ₁₀	22.5	<u>Net Conversion</u>	5.3	<u>Net Conversion</u>
<u>Others</u>	6	1.3	0.2	0.4

As can be seen in Table 4 above, the recycling of pentenes increases hexene selectivity and provides for a reduced pentene make. In these tests, between 26 and 38% of the pentenes in the feed were converted when approximately 10% pentene was present in the feed. The selectivity increase is most significant when greater amounts of pentenes are produced. Note further, however, when using catalyst systems with high isomerization activity, the amounts of pentenes formed are so large as to render recycle impractical and very costly. This effect is most significant when using low isomerization activity catalysts.

The disclosures of all patents and publications (including published patent applications) are hereby incorporated by reference to the same extent as if each patent and publication were incorporated individually by reference.

It is to be understood, however, that the scope of the present invention is not to be limited to the specific embodiments described above. The invention may be practiced other than as particularly described and still be within the scope of the accompanying claims.

WHAT IS CLAIMED IS:

1. A process for converting a first linear alpha olefin having carbon number n to ethylene and a second linear alpha-olefin having carbon number $2n - 2$, wherein $2n - 2$ is greater than n , said process comprising:

- (a) subjecting a feed including said first linear alpha-olefin having carbon number n to catalytic metathesis under conditions and with a metathesis catalyst that minimizes isomerization of said feed to produce an effluent comprising ethylene and an internal olefin having carbon number $2n - 2$;
- (b) fractionating said effluent from step (a) into a first stream containing said internal olefin having carbon number $2n - 2$ and a second stream including unreacted first linear alpha olefin and an internal olefin having carbon number $2n - 3$;
- (c) subjecting said first stream including said internal olefin having carbon number $2n - 2$ to isomerization to convert said internal olefin having carbon number $2n - 2$ to said second alpha olefin having carbon number $2n - 2$;
- (d) subjecting said second stream of step (b) to metathesis; and
- (e) combining the effluent from step (d) with said effluent of step (a); and
- (f) subjecting said effluent from step (d) and effluent of step (a) to fractionation in step (b).

2. The process of claim 1 wherein said second stream containing said unreacted first alpha olefin and said internal olefin having carbon number $2n - 3$ is subjected to additional metathesis by recycling said second stream to step (a).

3. The process of Claim 1 wherein the first linear alpha olefin is butene-1, the second linear alpha olefin is hexene-1 and said internal olefin having carbon number $2n - 2$ is a C_6 internal olefin, and said linear internal olefin having carbon number $2n-3$ is pentene.

4. The process of Claim 3 wherein the feed is comprised of at least 90% of butene-1.

5. The process of Claim 3 wherein said catalyst is selected from the group consisting of tungsten oxide, molybdenum oxide, rhenium oxide, and mixtures thereof.

6. The process of Claim 5 wherein said catalyst is tungsten oxide.
7. The process of Claim 6 wherein said catalyst is supported tungsten oxide.
8. The process of Claim 1 wherein the metathesis is effected at a WHSV of from about 3 to about 200.
9. The process of Claim 8 wherein said metathesis is effected at a WHSV of from about 6 to about 40.
10. The process of Claim 1 wherein said metathesis is effected at a pressure of from about 10 psig to about 600 psig.
11. The process of Claim 10 wherein said metathesis is effected at a pressure of from about 30 psig to about 100 psig.
12. The process of claim 1 wherein said metathesis is effected at a temperature from 250° to 400° C
13. The process of Claim 6 wherein said catalyst is supported on a silica support.
14. The process of Claim 3 wherein said C₆ internal olefin is isomerized to hexene-1 at a WHSV of from about 3 to about 200.
15. The process of Claim 14 wherein said C₆ internal olefin is isomerized to hexene-1 at a pressure of from about 2 bar to about 40 bar.
16. The process of Claim 15 wherein said C₆ internal olefin is isomerized to hexene-1 at a pressure of from about 3 bar to about 10 bar.

17. The process of Claim 3 wherein said C₆ internal olefin is isomerized to hexene-1 at a temperature of from about 40°C to about 400°C.
18. The process of Claim 17 wherein said C₆ internal olefin is isomerized to hexene-1 at a temperature of from about 250°C to about 350°C.
19. A process for producing ethylene and hexene-1 from a mixed C₄ feed containing butene-1, butene-2, butadiene, and isobutylene, comprising:
- (a) removing butadiene;
 - (b) hydroisomerizing butene-1 to butene-2;
 - (c) separating isobutylene from other butenes;
 - (d) isomerizing butene-2 from step (b) to butene-1;
 - (e) subjecting butene-1 from step (d) to catalytic metathesis under conditions and in the presence of a metathesis catalyst that minimizes isomerization to produce an effluent including ethylene, pentene-2, and hexene-3, and wherein said pentene-2 is recycled to said butene-1 produced from step (d) to be subjected to said catalytic metathesis;
 - (f) fractionating the effluent; and
 - (g) isomerizing hexene-3 from step (e) to produce hexene-1.
20. The process of Claim 1 wherein said catalyst does not include a promoter.
21. The process of Claim 19 wherein said catalyst does not include a promoter.

INTERNATIONAL SEARCH REPORT

International Application No

PCT/US 02/02352

A. CLASSIFICATION OF SUBJECT MATTER

IPC 7 C07C11/04 C07C11/107 C07C6/04 C07C5/25

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 C07C

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

EPO-Internal

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	US 3 776 974 A (GAUTIER P ET AL) 4 December 1973 (1973-12-04) ---	5-13
A	GB 1 471 151 A (SHELL INTERNATIONALE RESEARCH MAATSCHAPPIJ) 21 April 1977 (1977-04-21) ---	
X	US 4 368 345 A (DICKINSON KATHLEEN F) 11 January 1983 (1983-01-11) claims -----	

☐ Further documents are listed in the continuation of box C.

Patent family members are listed in annex.

* Special categories of cited documents :

A document defining the general state of the art which is not considered to be of particular relevance

E earlier document but published on or after the international filing date

L document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)

O document referring to an oral disclosure, use, exhibition or other means

P document published prior to the international filing date but later than the priority date claimed

T later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

X document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

Y document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.

A document member of the same patent family

Date of the actual completion of the international search

6 June 2002

Date of mailing of the international search report

19/06/2002

Name and mailing address of the ISA

European Patent Office, P.B. 5818 Patentlaan 2
NL - 2280 HV Rijswijk
Tel. (+31-70) 340-2040, Tx. 31 651 epo nl,
Fax: (+31-70) 340-3016

Authorized officer

Van Geyt, J

INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

PCT/US 02/02352

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
US 3776974	A	04-12-1973	GB 1357632 A	26-06-1974
			BE 779237 A2	11-08-1972
			CA 982153 A1	20-01-1976
			DE 2208058 A1	07-09-1972
			FR 2126245 A5	06-10-1972
			IT 947758 B	30-05-1973
			JP 58002207 B	14-01-1983
			NL 7202235 A ,B,	24-08-1972
GB 1471151	A	21-04-1977	BE 815329 A2	21-11-1974
			CA 1034968 A1	18-07-1978
			DE 2425685 A1	19-12-1974
			FR 2231641 A1	27-12-1974
			IT 1012814 B	10-03-1977
			JP 1207424 C	11-05-1984
			JP 50019703 A	01-03-1975
			JP 58043373 B	27-09-1983
US 4368345	A	11-01-1983	NL 7407100 A ,B,	03-12-1974
			DE 56013 T1	28-04-1983
			EP 0056013 A2	14-07-1982
			JP 58126818 A	28-07-1983